

Supporting Information

Chemical-Vapor-Deposition-Synthesized Two-Dimensional Non-Stoichiometric Copper Selenide (β -Cu_{2-x}Se) for Ultra-Fast Tetracycline Hydrochloride Degradation under Solar Light

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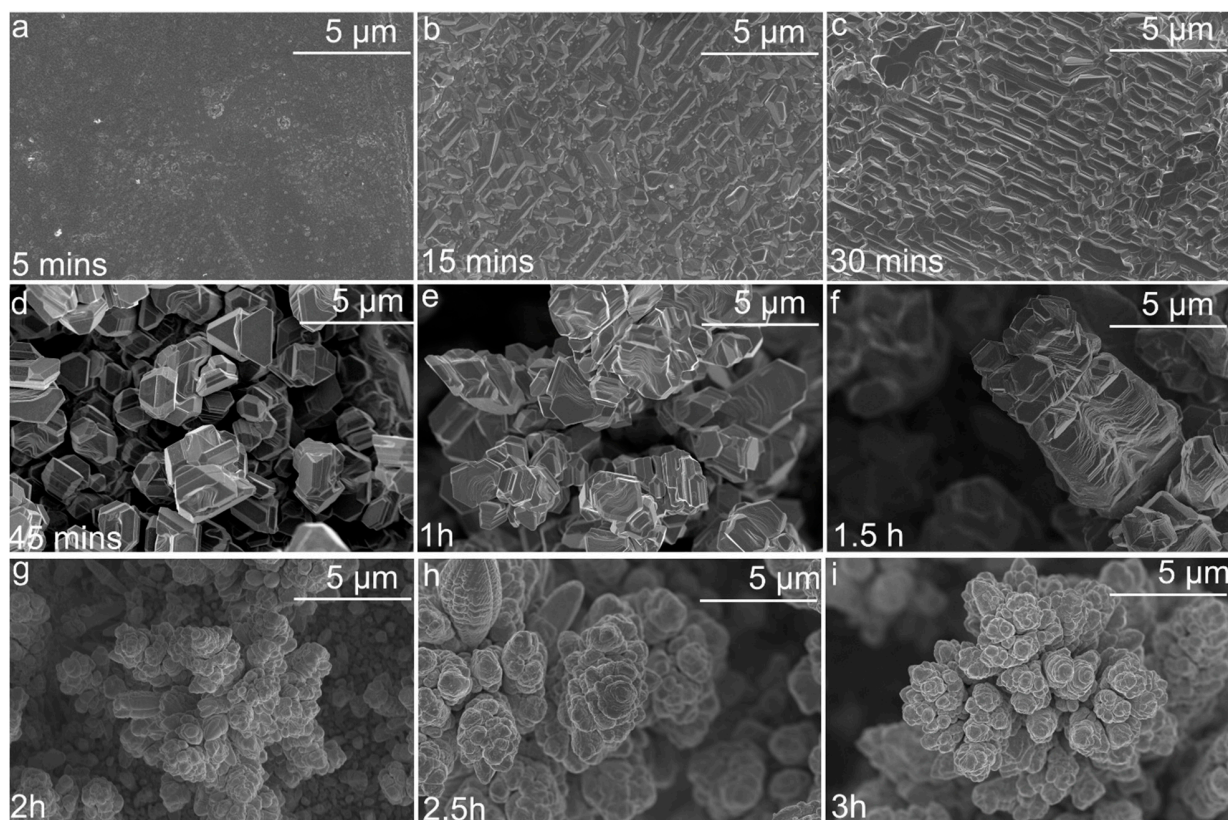


Figure S1. High-resolution scanning electron microscopy (HRSEM) shows microstructural evolution and copper selenide ($\beta\text{-Cu}_{2-x}\text{Se}$) formation at different reaction times (5mins to 3h) between etched Cu foil and elemental Se.

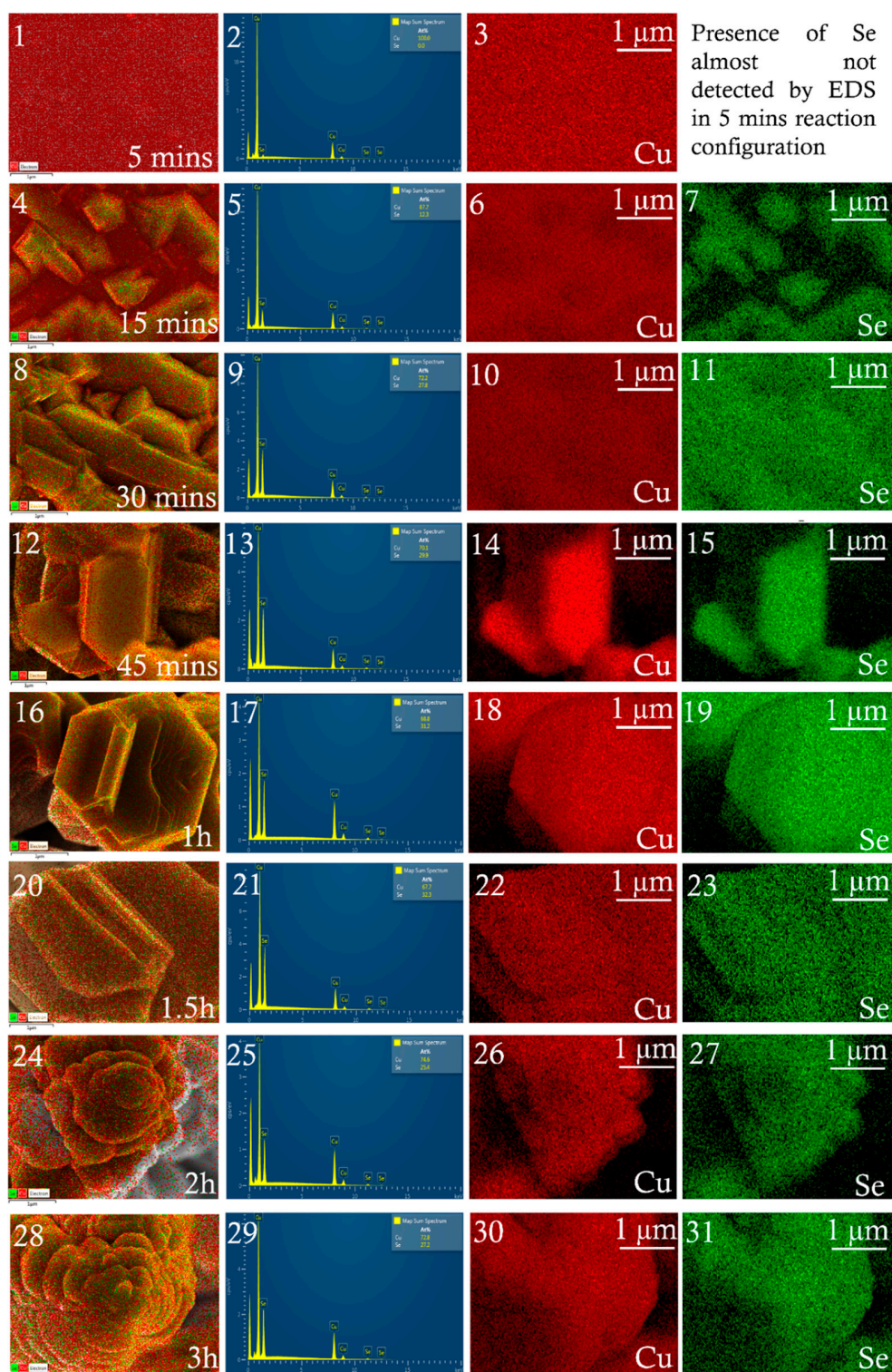


Figure S2. Ex-situ High-resolution scanning electron microscopy (HRSEM) images, Electron Dispersive Spectroscopy (EDS), and elemental mapping of Cu and Se at different reaction times indicate the formation and evolution of $\beta\text{-Cu}_{2-x}\text{Se}$.

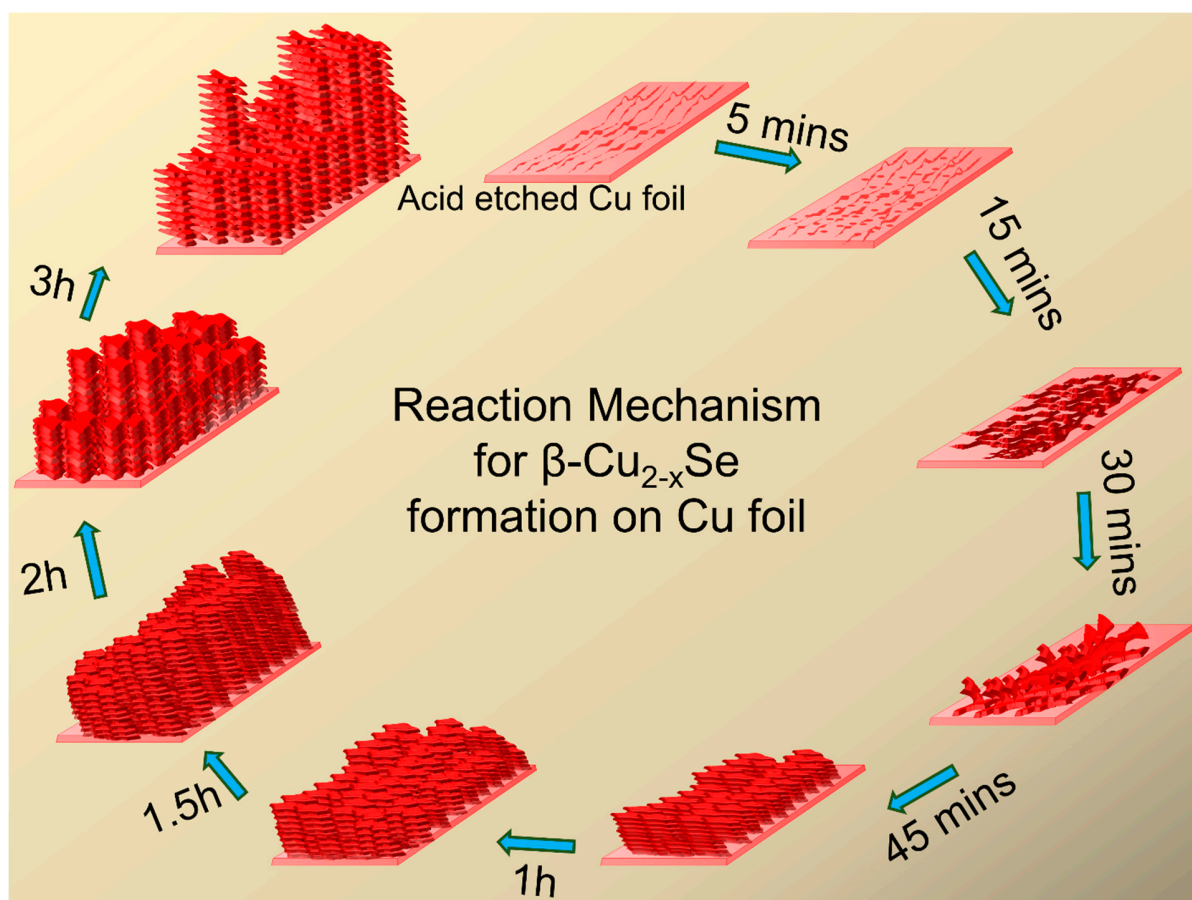


Figure S3. Proposed reaction mechanism based on the microstructural evolution and CVD-synthesis of $\beta\text{-Cu}_{2-x}\text{Se}$ on Cu foil.

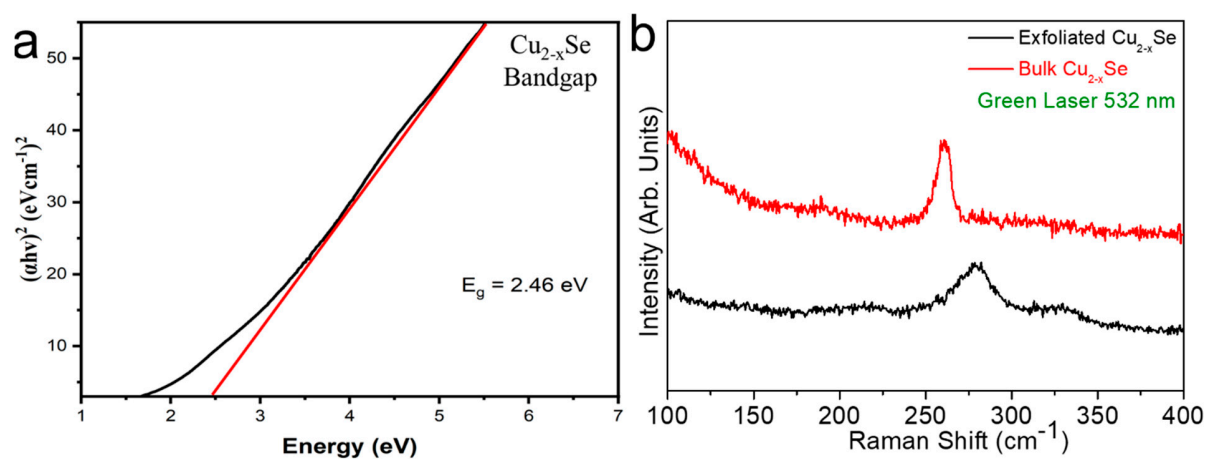


Figure S4. (a) Band gap and (b) Raman measurements of $\beta\text{-Cu}_{2-x}\text{Se}$.

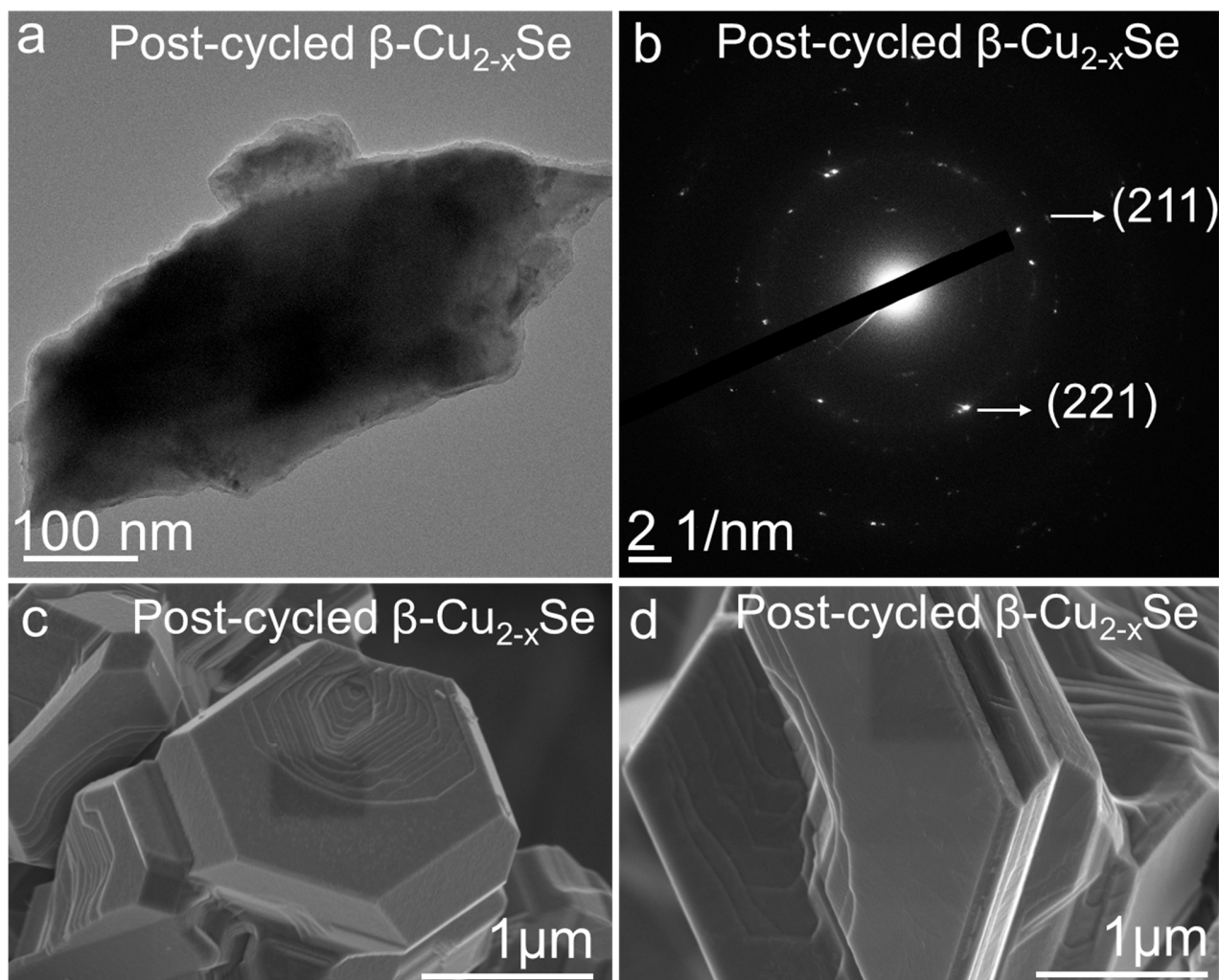


Figure S5. Post-mortem studies showing (a) High-resolution transmission electron microscopy (HRTEM) image of $\beta\text{-Cu}_{2-x}\text{Se}$ after photocatalytic degradation (4 cycles); (b) Selected Area Diffraction (SAED) of post-cycled $\beta\text{-Cu}_{2-x}\text{Se}$; (c) and (d) High-resolution scanning electron microscopy (HRSEM) images of bulk $\beta\text{-Cu}_{2-x}\text{Se}$ indicating that the material is very stable and does not undergo extreme microstructural changes after photocatalytic cycling (4 cycles).

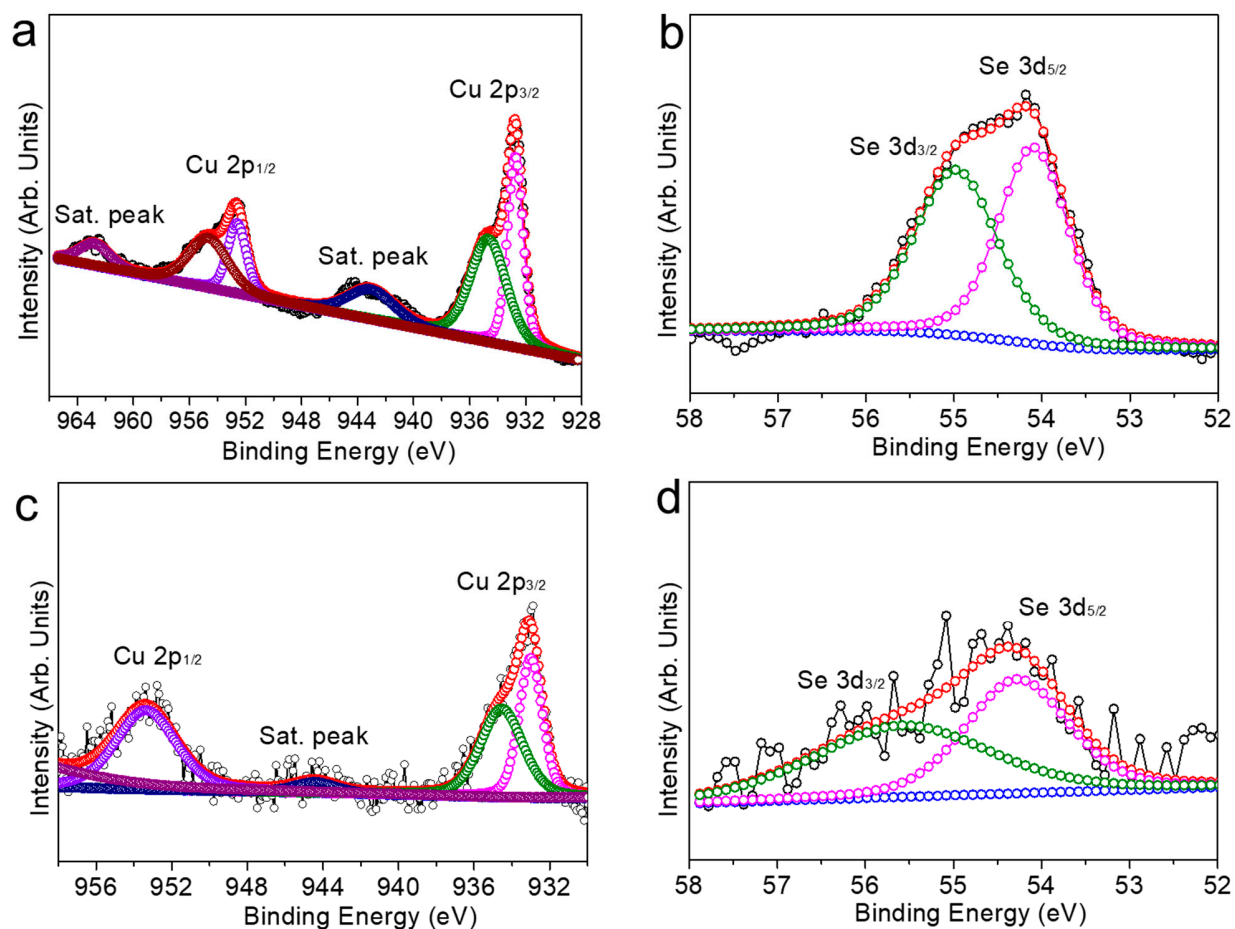


Figure S6. (a) and (b) are the X-Ray Photoelectron Spectroscopy (XPS) measurements of uncycled $\beta\text{-Cu}_{2-x}\text{Se}$, while (c) and (d) are the XPS) measurements of cycled $\beta\text{-Cu}_{2-x}\text{Se}$ under simulated solar light.

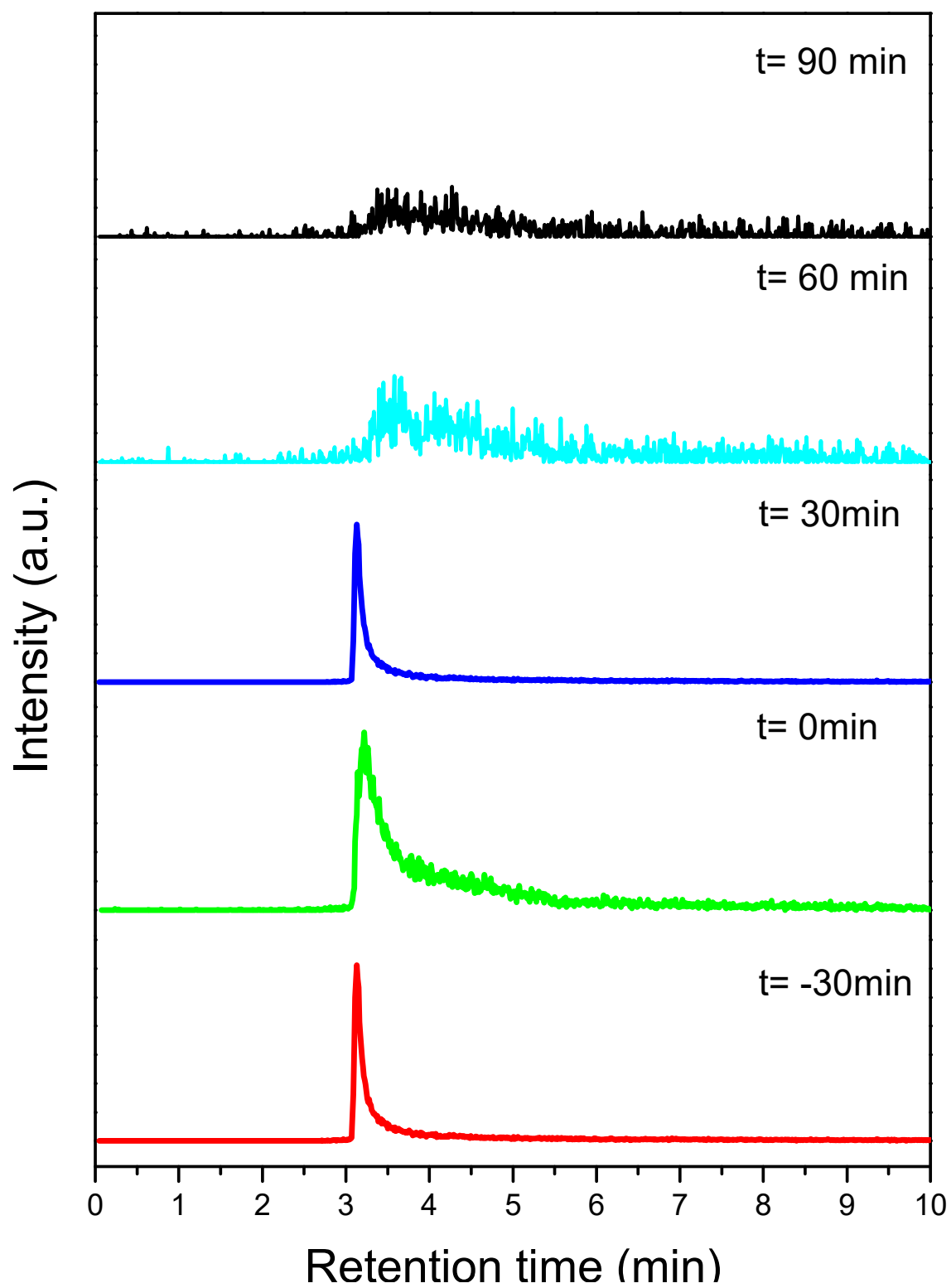


Figure S7. LC-MS of TCH degradation over β -Cu_{2-x}Se under solar light irradiation.

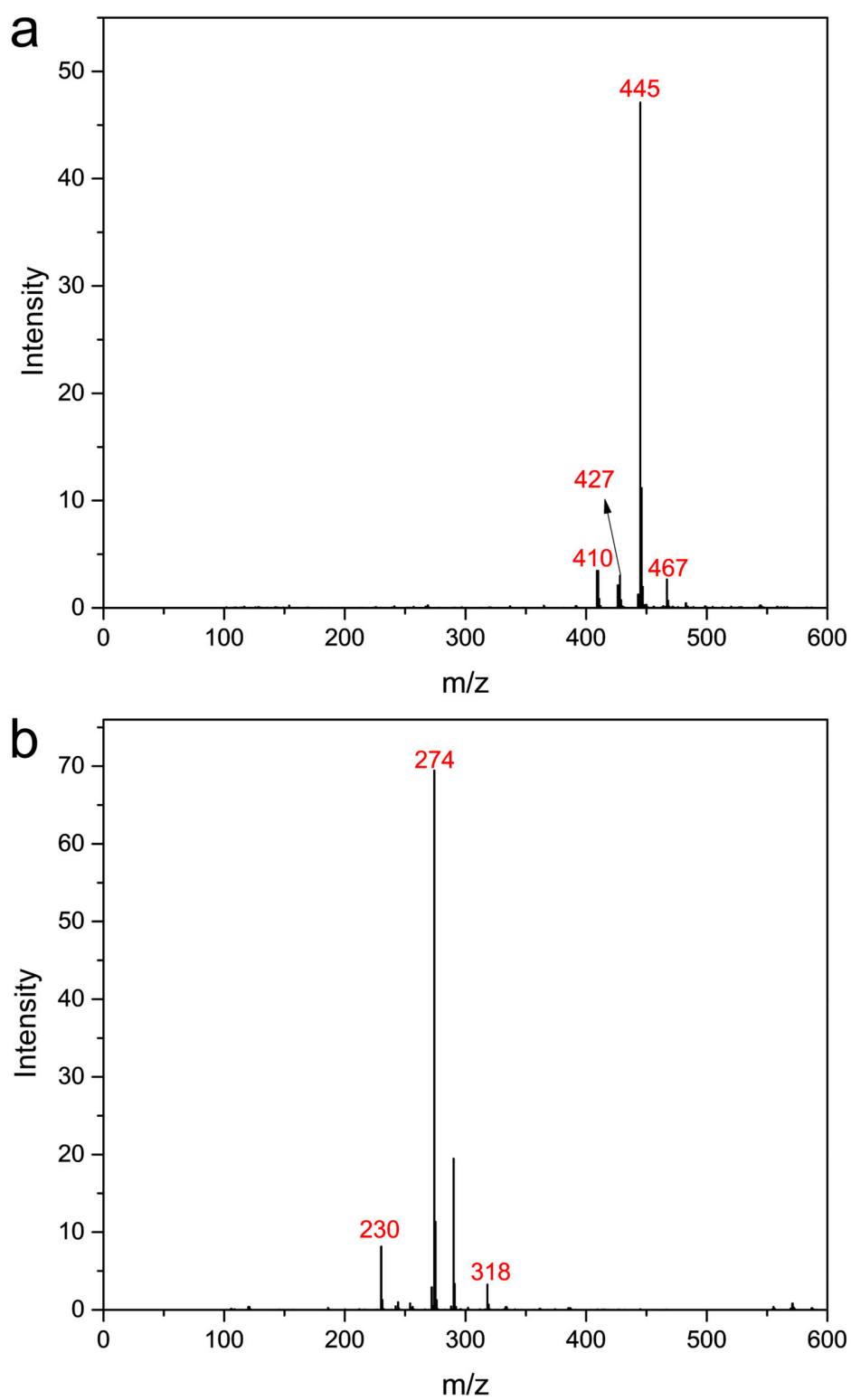


Figure S8. The mass spectra of TC-HCl transformation products: m/z = 467, 445, 427, 410, 318, 274, 230.

Table S1. Comparison of TC-HCl photodegradation using different photocatalysts

Catalyst	Concentration of TC-HCl	Catalyst dosage	TC-HCl /Catalyst dosage	Degradation time	Efficiency	Reaction rate	Ref.
β-Cu_{2-x}Se	20 mg/L	0.4 g/L	50 mg/g	90 min	98.35 %	0.0314 min⁻¹	Present Work
TiS ₂	20 mg/L	1 g/L	20 mg/g	300 min	95.43 %	0.0174 min ⁻¹	[1]
Sn ₃ O ₄ /g-C ₃ N ₄	10 mg/L	0.5 g/L	20 mg/g	120 min	72.2 %	0.0108 min ⁻¹	[2]
CdS/Ti ₃ 2-oxo-cluster	50 mg/L	0.2 g/L	250 mg/g	60 min	96.3 %	0.06135 min ⁻¹	[3]
Bi ₂₄ O ₃₁ Br ₁₀	20 mg /L	0.3 g/L	67 mg/g	90 min	95 %	0.031 min ⁻¹	[4]
WO ₃ /g-C ₃ N ₄ /Bi ₂ O ₃	10 mg/L	1 g/L	10 mg/g	60 min	80.2 %	0.0236 min ⁻¹	[5]
Ag/AgIn ₅ S ₈	10 mg/L	0.3 g/L	33 mg/g	120 min	95.3 %	0.023 min ⁻¹	[6]
BiOCl	10 mg/L	0.5 g/L	20 mg/g	90 min	71.8 %	0.0139 min ⁻¹	[7]
WO ₃ /g-C ₃ N ₄	20 mg/L	1 g/L	20 mg/g	60 min	90.54 %	0.0378 min ⁻¹	[8]
CNFs/g-C ₃ N ₄ /BiOBr	20 mg/L	3 g/L	6.7 mg/g	120 min	86.1 %	0.015 min ⁻¹	[9]
CDs/MoO ₃ /g-C ₃ N ₄	20 mg/L	0.6 g/L	33.3 mg/g	90 min	88.4 %	0.0231 min ⁻¹	[10]
ZnIn ₂ S ₄ /MoO ₃	30 mg/L	0.15 g/L	200 mg/g	90 min	94.5 %	0.0292 min ⁻¹	[11]

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